Name:	Date:
Notes: Trends of the	e Periodic Table
Where are metals located?	
Where are nonmetals located?	
What do we call the elements lying along the dark,	jagged line?
Why are metalloids unique?	·
Donating and Gaining Electrons	Group 16:
Group 1:	Group 17:
Group 2:	Group 18:
Group 1:	Group 17:
Group 2:	Group 18:
What do we call the metals in the d-block?	
Label the periodic table below with the predicted	charge of groups 1, 2, and 13-18.
What kinds of elements are in ionic compounds?	H He Li Be B C N O F Ne
What kinds of elements are in covalent compound	Rb Sr Y Zr Nb Mo Tc Ru Rh Pd Ag Cd In Sn Sb Te I Xe
Shade the periodic table to the right showing the s, p, d, and f regions.	Fr Ra Ac Rf Db Sg Bh Hs Mt Ce Pr Nd Pm Sm Eu Gd Tb Dy Ho Er Tm Yb Lu Th Pa U Np Pu Am Cm Bk Cf Es Fm Md No Lr
Define <i>atomic radius</i> :	
What makes atoms larger as you go down a group?	?
What else affects the size of the atomic radius of a	toms?
Why does the size of atoms shrink as you move to	the right along a period?
What element has the largest atomic radius?	
How do the atomic radii of elements change as you	u move through groups and periods?
1.	
2	

When do atoms b	ecor	me i	ions? _																			
Define <i>ionic radiu</i>	s:																				=	
What happens to	the	elec	trons	of m	eta	<i>ls</i> ?															_	
What do we call p	ositi	ively	ı charg	ied id	ons	?															- _^	
How does a metal	ion	cor	npare	to th	ne d	rigi	nal a	ator	m?												- 💢	7
What happens to	the	elec	trons	of nc	onn	neto	als?														_	
What do we call <i>n</i>	ega	tive	ly char	ged	ion	s? _																
What do we call <i>n</i> How does a nonm	etal	l ion	comp	are t	to t	he d	origi	nal	ato	m?											-	7
Define <i>ionization</i>																					_ `	
Why does it take r	nor	e en	ergy to	o rer	nov	⁄e a	n ele	ectr	on	fror	m a	no	nme	etal	tha	n fr	om	a m	etal	?		
																					- <u>,</u>	
What element has	s the	e gr	eatest	ioni	zati	ion	enei	rgy	?												->>	7
How does the ioni																						
1																					_	
2																					_	
Define <i>electroneg</i>	ativ	ity:																		\leftarrow	. \((-	-)
																		^	H)—(11	1
What element has	s the	e gr	eatest	elec	tro	neg	ativ	ity?	?								`	\mathcal{N}	/ (+	+) (H	
How does the elec																	ups	and	peri	iods?		
1																					_	
2																					_	
	Н																He					
	Li	Ве										В	С	N	0	F	Ne					
	Na	-										ΑI	Si	Р	S	CI						
	K	Ca	Sc Ti	V	Cr	Mn	Fe (Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr					

Н																	He
Li	Ве											В	С	N	0	F	Ne
Na	Mg											ΑI	Si	Р	s	CI	Ar
K	Ca	Sc	Ti	٧	Cr	Mn	Fe	Со	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Rb	Sr	Υ	Zr	Nb	Мо	Тс	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Хе
Cs	Ва	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	TI	Pb	Bi	Ро	At	Rn
Fr	Ra	Ac	Rf	Db	Sg	Bh	Hs	Mt									
				Се	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Но	Er	Tm	Yb	Lu
				Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr