Name:	Date:
Notes: The Periodic	Table of the Elements
Who developed the first accepted Periodic Tab	ole?
What helped validate this periodic table and the	ne arrangement of elements it presented?
Describe <i>atomic number</i> :	
Describe <i>atomic mass</i> :	
Where are metals located?	
Where are nonmetals located?	
What happens to the electrons of <i>metals</i> ?	
What kind of charge do metals have?	
What do we call positively charged ions?	
Define oxidation:	
What happens to the electrons of <i>nonmetals</i> ?	
What kind of charge do nonmetals have?	
What do we call negatively charged ions?	
Define reduction :	
Why types of elements do metals react with?	
Donating and Gaining Electrons	Group 16:
Group 1:	Group 17:
Group 2:	Group 18:
What do we call the metals in the d-block?	
What do we call the elements lying along the c	dark, jagged line?
Why are metalloids unique?	
What 3 elements compose the Iron Triad?	
Which elements are filling their s and p subleven	els?
Which elements are filling their d and f subleve	els?
GROUP NAMES:	
Group 1:	Group 17:
Group 2:	Group 18:

What do we call colum	ins on the periodic ta	able <i>beside groups</i> ?			
What do elements in a group have in common?					
Why do these element	ts have similar prope	rties?			
Which group on the pe	eriodic table undergo	es the fewest reaction	ns?		
What does <i>inert</i> mean	?				
Label the periodic table representative elemen	_	e predicted charge of	the ions made from		
Use your periodic tabl of the following ions:	e to predict the char	rges H	ВС	N O F N	He Ne
Na	Cl Ca	Na Mg	V Cr Mn Fe Co Ni Cu Zn Ga Ge		
	P F	Dh Sr V 7r	Nb Mo Tc Ru Rh Pd Ag Cd In Sr		_
		CS Ba La Hi	Ta W Re Os Ir Pt Au Hg TI Pt Db Sg Bh Hs Mt	Bi Po At R	₹n
	K Mg			T= T= T= T=	
Be	I Ar		Ce Pr Nd Pm Sm Eu Gd Tb Dy Ho Th Pa U Np Pu Am Cm Bk Cf Es		
What are the charges	of the following tran	nsition metals?			
Palladium (II)	Lead (IV)	Cobal	lt (III)		
Vanadium (IV)	Platinum (II) _	Mang	ganese (III)		
Why do transition elements always have a positive charge?					
What kind of elements	s are in an <i>ionic comp</i>	oound?			
What kind of elements are in a <i>covalent compound</i> ?					
What can nonmetals do that allows covalent compounds to form?					
		VF			
1 The formula above s	shows the ratio of an	XF	· V to fluoring in an ionic s	alt	
Which group below is			, X, to fluorine in an ionic s	ait.	
<u> </u>	B lithium	C arsenic	D krypton		
		XCl ₃			
2. The formula above s	shows the ratio of an	•	c, X, to chlorine in a covaler	nt	
molecule. Which elem	nent below is most lik	cely the identity of Ele	ement X?		
A sodium	B neon	C phosphorous	D copper		