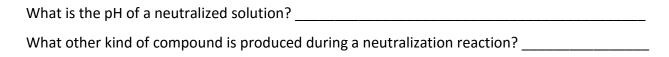
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Name:	Date:	
Notes: Double Displacement Reactions		
What happens to the elec	ctrons of <i>metals</i> ?	
What type of charge do n	netal ions have?	
Define oxidation :		
What happens to the elec	ctrons of <i>nonmetals</i> ?	
What type of charge do <i>n</i>	nonmetal ions have?	
Define reduction :		
	and reduction?	
How would a chemical re	action occur that was not a redox reaction?	
Why would this type of re	eaction not be considered an oxidation-reduction reaction?	
Describe <i>double displace</i>	ment reactions:	
	reaction:	
What is a precipitate ?		
What is true of the <i>reacto</i>	ants in a precipitation reaction?	
What happens to ionic co	ompounds when they dissolve in water?	
	solutions mix?	
	utralization reaction:	
What liquid is produced v	when an acid neutralizes base?	

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WHAT ARE THE KEY CHARACTERISTICS OF A PRECIPITATION REACTION?

WHAT ARE THE KEY CHARACTERISTICS OF AN ACID-BASE NEUTRALIZATION REACTION?

For each of the double displacement reactions on your notes, determine if the reaction is a precipitation reaction or a neutralization reaction.

$$2NaOH_{(aq)} + H_2CO_{3(aq)} \rightarrow 2H_2O_{(I)} + Na_2CO_{3(aq)} \underline{\hspace{2cm}} \\ Mg(NO_3)_{2(aq)} + Li_2SO_{4(aq)} \rightarrow MgSO_{4(s)} + 2LiNO_{3(aq)} \underline{\hspace{2cm}} \\ Pb(C_2H_3O_2)_2 + NaCI \rightarrow NaC_2H_3O_2 + PbCI_2 \underline{\hspace{2cm}} \\ 2HCI + Ca(OH)_2 \rightarrow CaCI_2 + 2H_2O \underline{\hspace{2cm}} \\ Hg_2(NO_3)_{2(aq)} + 2KBr_{(aq)} \rightarrow Hg_2Br_{2(s)} + 2KNO_{3(aq)} \underline{\hspace{2cm}} \\ 2LiOH + 2H_2SO_3 \rightarrow 2H_2O + Li_2SO_3 \underline{\hspace{2cm}} \\ Sr(OH)_2 + 2HCI \rightarrow SrCI_2 + 2H_2O \underline{\hspace{2cm}} \\ CaSO_4 + SrBr_2 \rightarrow CaBr_2 + SrSO_4 \underline{\hspace{2cm}} \\ AgNO_3 + KCI \rightarrow AgCI + KNO_3 \underline{\hspace{2cm}} \\ Ba(OH)_{2(aq)} + 2HI_{(aq)} \rightarrow 2H_2O_{(I)} + BaI_{2(aq)} \underline{\hspace{2cm}} \\ Fe(C_2H_3O_2)_2 + 2NaOH \rightarrow 2NaC_2H_3O_2 + Fe(OH)_2 \underline{\hspace{2cm}} \\ H_3PO_4 + 3KOH + \rightarrow 3H_2O + K_3PO_4 \underline{\hspace{2cm}} \\ \nearrow NH_3(aq) + HNO_3(aq) \rightarrow NH_4NO_3(aq) \underline{\hspace{2cm}}$$